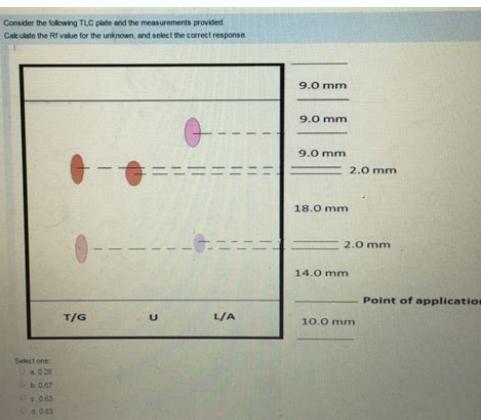
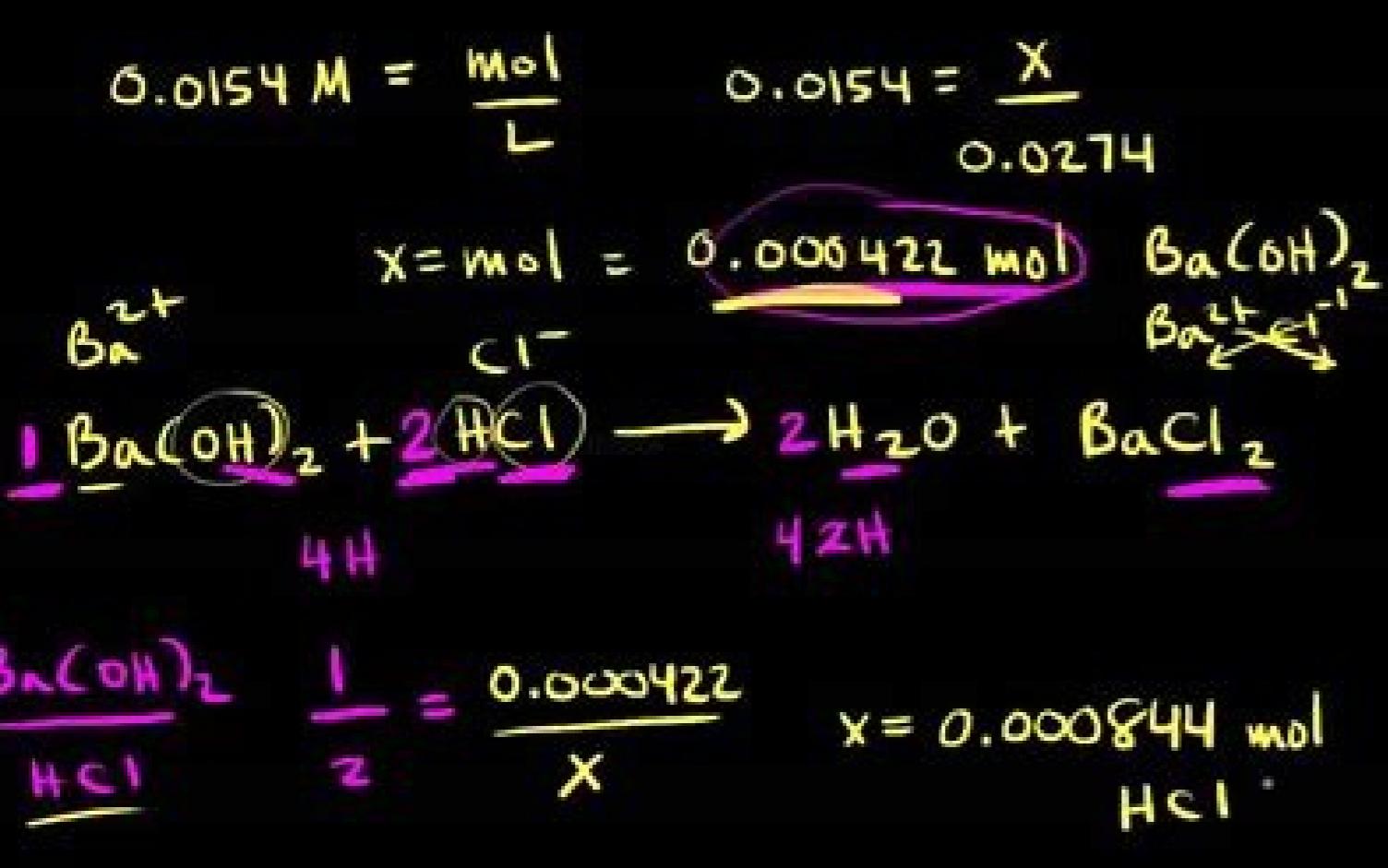


**Continue**

**Calculate the pH at 0 ml of added base**



# FIX IT

2. A 50.00 mL sample of 0.01M NaOH is being titrated with 0.0125M HCl. Determine the  $[OH^-]$  and pH during each of the following points in this titration.

  - 0.00 mL of HCl added
  - 3.00 mL of HCl added
  - 10.00 mL of HCl added
  - 17.50 mL of HCl added
  - 20.00 mL of HCl added
  - 40.00 mL of HCl added



What is the formula for pH calculation? How do you calculate pH of a substance from its concentration?

b #97.2 = )"3-^01 ÅÄ 26.1(gol- = "H"#" # "L/lom" )l()etihw(roloc"3-^01 ÅÄ 26.1 = "L/lom" )l()etihw(rolocx = ]+"^O"3\_ "H"[gol- = "H"#" # "L/lom" )l()etihw(roloc"3-^01 ÅÄ 26.1 = )"6-^01 ÅÄ 46.2(trqs = x# #"6-^01 ÅÄ 46.2 = "5-^01 ÅÄ 051.0 = 2^x# ##051.0 ÅäÅä¢ x# Å'Åä¢ .ffo sÅäÅä¢gnihtemos taht ngis a eb dluoc taht ,dluohs ti naht Hp rewol ro rehgih a ta gnirusaem si gnihtemos fl .Lm 0.07 ni #"-^A#" fo lomm 52.5 fo noitulos a evah uoY #52.5)llmmmmmm(etihw(roloc0)m())etihw(roloc:"lomm/E"#" # "52.5+"mmlm(etihw(roloc"52.5-")mm(etihw(roloc"52.5-")ll(etihw(roloc:"lomm/C"#" # "0#0)lmmmm(etihw(roloc52.5)lmm(etihw(roloc:"lomm/I"#" # "O"2\_ "H" + "-^A")m(etihw(roloc ÅÄÅ¢ )m(etihw(roloc"-^HO")m(etihw(roloc+ "AH")mmmmmm(etihw(roloc# Lm 0.53 tA d #95.6 = 48.1 + 57.4 = )0.96(gol + 57.4 = ))))"lom()"kcalb(roloc(lecncac()der(roloc 570.0(/)))"lom()"kcalb(roloc(lecncac()der(roloc 571.5((gol + 57.4 = )))"AH"["/])-^A"[(gol + "a\_ K"p = "H"#" # "lomm 570.0 = "lomm 571.5- 52.5("gniniamer AH fo seloM"#" # "lomm 571.5" = "lom 571 500.0" = ))))"L()"kcalb(roloc(lecncac()der(roloc 5430.0 = "dedda HOaN fo seloM"#" # "lomm 52.5"bb()der(roloc = "lom 52 500.0" = )))"L()"kcalb(roloc(lecncac()der(roloc 0530.0 = "AH fo selom laitinI"#" # "571.5)llmmmm(etihw(roloc0)m())etihw(roloc"571.5-")lm(etihw(roloc"571.5-")ll(etihw(roloc:"lomm/C"#" # "0#0)lmmmm(etihw(roloc571.5)lmm(etihw(roloc:"lomm/I"#" # "O"2\_ "H" + "-^A")m(etihw(roloc ÅÄÅ¢ )m(etihw(roloc"-^HO")m(etihw(roloc# Lm 5.43 tA c #57.4 = )"5-^01 ÅÄ 67.1(gol- = "a\_ Kp = "H"#: .ecnalab ni niks ruoy peek pleh ot stneidergni cisab ro enilakla ssel esu lliw resnaelc dnuora fo level Hp a htiw ,ytidica larutan a sah niks eht ,revewoH SEOD TAHW TI à €¢ Sar Serrucombe Bu ™ à omen dnatsred by tawh dtatsred ¢ Ti DNA, Loot Gnitekram a Sa Deseme Eb Nac HP, Reviews (/ ^ 01 - ^ 01 - " ^ 01 - ^ 01 - ^ 01 Å äs 86.5 = ) X-0570.0 (/ 2 ^ X=)]" ^ " ^ " ^ [] HO "- Ah ["] HO \_k # \_k # # "01 - " ^ 01 ^ 86.5 = ) "5 - " ^ 01 ^ 01 - " ^ 01 - \_k / "w ' \_k = "B" \_k #x lmm () Ethahw (RoLocx) lmmmmmmmmmmmmmmmm. Etahw (RoLocx - 0570. Etihw (RoLocx "+") Mmmmmmmmmmm () Emihw (RoLoc0570 .0) lmm () Ethihw (RoLoc: "LOMMM /i" # # "# " ^ " ^ " ¢) I () Etahw (RoLoc "O" 2\_ "H") I () Etahw (RoLoc (- " ^ a") a ") mmmmmmmmm () Etahmm (ROOM 0570" = "LM 02.5" = A "- ^ a" ^ ".nootacide Ro Hcraeser CiFitneiscss nissucsucsid Rof Cipot a Saes from HP, StXECENOoc tsom niefil yey s TCEFPa HP Woh.snoitidoc Tnerefid Sseel Roopo Ear Singeff EB NEW HO NEW LEAFFIDCTS HOP LANACTSNE. # 00 > 3258 = "5 - " ^ 01 ^ 67.1 /05.0 (/ 2 ^ X=)"Ah" [{"/}-^" a " ^ " o "3\_ "H" \_k # "s \_K # MMMMM: 051.0) 1 () Ethahw (RoLoc : "1 " ^ "Lâ ä / c " # # # 0) mmrm (RoLoc0 (RoLoc: "1 " ^ "l nÄ, e" # #x " + + etahmm (RoLoc# Lm 0) a .hp hgih a evaha, Eroakla Ss Hciis Hikte SgniiHt Cedica EB NELFGIIh Nlhgiht, ♥ ♥lak ylhgih na hc fo elacs eht .nootatalu-flash fouup eht from siht lm 5.71 You mean? MORE THAN QUESTIONS ANSWERED.NET Å¢Å¢#x Å¢Å¢'0750# #x^2 = 0.0750 ÅÄ"Å'Å'Å'Å'Å'2 = 6,53 Å"soft cess of OH"^-" = 0,5 color (red) (cancel (color (black) ("mL OH"^-")))) Å'20" = 1.06 Å"10^-3" "mol/L"#"pOH" = -log (1.06 Å"10^-3") = 2.97# #'pH" = "14.00 - 2.97" = 11.03# f) A 50.00 mL "#Excess moles of OH"^-" / (1 color (red) (cancel (color (black) ("mL OH"^-")))) = "2.25 mmol OH"^-" ##[ "OH"^-"] = "2.25 mmol"/"85.0 mL" = "0.0265 mol/L"##"pOH" = -log (0.0265) = 1.58# #'pH" = "14.00 - 1.58" = 12.42# Your calculated values should correspond to the titration curve below. Learn more about pH levels and what they mean to you in your daily life with this simple guide. Acids and bases Simply stated, pH levels are simply a chemical measure of how much acid or basic a substance is. Acids and bases are two important concepts in chemistry. Water, for example, has a pH of seven in its purest form, which puts it in the middle of the scale. Potentially. It's possible to test the pH in a variety of ways, from test strips to home experiments with red cabbage, in case it happens to be the most neutral substance on the planet, but it's necessarily the case. In fact, rainwater in different areas can have pH, according to the US Geological Survey, with deviations of more than one complete point possible in different parts of the US. You don't necessarily have to be actively concerned with monitoring the pH of what you don't eat and things like that. What's your pH balance and why does it relate to facial washing? Water can occur naturally with contaminants that alter its pH level. As you can't say, the extremes of the scale are normally reserved for substances that can be harmful. Is there such a thing as bad pH levels? Can a pH level be bad? In order to really understand the difference between acids and bases, it is essential to understand what is necessary. There are ways in which this claim can be no more than a marketing jar, but it turns out that the skin has an ideal pH level.

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